The Mole and Molar Mass

Question One:

1 mol of $CaCl_2$ = 6.02×10^{23} _____ of $CaCl_2$

Question Two:

1 mol of H_2O_2 = 6.02×10^{23} _____ of H_2O_2

Question Three:

1 mol of Cu = 6.02×10^{23} _____ of Cu

Question Four:

1 mol of H_2 = 6.02×10^{23} _____ of H_2

Question Five:

1 mol of NaF = 6.02×10^{23} _____ of NaF

Question Six:

1 mol of Fe = 6.02×10^{23} _____ of Fe

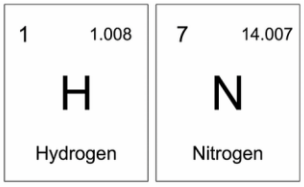
Question Seven:

1 mol of CO_2 = 6.02×10^{23} _____ of CO_2

Question Eight:

1 mol of O_2 = 6.02×10^{23} _____ of O_2

Example One: Determine the molar mass for ammonia, NH_3 to one decimal place.



Example Two: Determine the molar mass of Aluminum Sulfate, $Al_2(SO_4)_3$ to one decimal place.